



Hydroflex™ Hydrogen Reference Electrode (Model ET070)

US Patent: 5 407 555



- No external hydrogen source required
- Pressures up to 10 bar
- Temperatures up to 210 °C
- Replaceable H₂ cartridges

Description

A self-contained hydrogen reference electrode with replaceable internal hydrogen source (one cartridge included).

Compatibility

Low impedance ensures the electrode will be suitable for use with most potentiometers, pH and ISE meters, and potentiostats.

Applications

You can use the Hydroflex electrode to calibrate your everyday Ag/AgCl or calomel electrodes, or use it as the reference electrode in situations where other electrodes might not be suitable. For example, Hydroflex is particularly suitable as a reference electrode in aqueous acid or alkali solutions, in fuel cell studies, and can be used at pressures up to 10 bar and temperatures of up to 210 °C.

Theory of Operation

Traditionally, a hydrogen reference electrode requires passing hydrogen gas over a platinized platinum sheet. Configuring the platinum plate and bubbler mechanism within an electrode body is not a trivial matter, and the need for a hydrogen gas cylinder creates issues with space, expense, and safety. No wonder that workers so often resort to other types of reference electrode!

The Hydroflex electrode has an internal, replaceable, cartridge (order part number ET071) that continuously generates a low volume hydrogen flow that passes through a platinized gas diffusion electrode, resulting in a compact reference electrode without the need for a separate hydrogen source.

Hydrogen flow rate of the replaceable cartridge can be adjusted to last between 1 – 12 months (approximately), but a setting of six months is recommended for most applications.

Specifications

Length	120 mm
Diameter/material of shaft:	8 mm / PTFE
Connector:	Socket for 2 mm pin (4mm adaptor included)
Replacement H ₂ cartridges:	ET071 (Set of 4)
<i>eDAQ Pty Ltd reserves the right to alter these specifications at any time.</i>	

Citations

- Fourier transform electrochemical impedance spectroscopic studies on platinum electrodes in an acidic medium. Jin-Bum Park, and Su-Moon Park. *Journal of Electroanalytical Chemistry*, in Press.
DOI: 10.1016/j.jelechem.2010.10.026
- Noble Metal-Free Hydrazine Fuel Cell Catalysts: EPOC Effect in Competing Chemical and Electrochemical Reaction Pathways. Jean Sanabria-Chinchilla, Koichiro Asazawa, Tomokazu Sakamoto, Koji Yamada, Hirohisa Tanaka, and Peter Strasser. *Journal of the American Chemical Society*, 133, 5425–5431, 2011. DOI: 10.1021/ja111160r
- Synthesis and Oxygen Reduction Electrocatalytic Property of Platinum Hollow and Platinum-on-Silver Nanoparticles. Zhenmeng Peng, Jianbo Wu, and Hong Yang. *Chemistry of Materials*, 22, 1098–1106, 2010.
DOI: 10.1021/cm902218j
- A novel cathode for alkaline fuel cells based on a porous silver membrane. F. Bidault, A. Kucernak. *Journal of Power Sources*, 195, 2549–2556, 2010.
DOI: 10.1016/j.jpowsour.2009.10.098
- Titanium nitride nanoparticles based electrocatalysts for proton exchange membrane fuel cells. Bharat Avasarala, Thomas Murray, Wenzhen Li and Pradeep Haldar. *Journal of Materials Chemistry*, 19, 1803–1805, 2009.
DOI: 10.1039/b819006b
- The influence of support and particle size on the platinum catalysed oxygen reduction reaction. Brian E. Hayden, Derek Pletcher, Jens-Peter Suchsland and Laura J. Williams. *Physical Chemistry Chemical Physics*, 11, 9141–9148, 2009. DOI: 10.1039/b910110a
- The influence of Pt particle size on the surface oxidation of titania supported platinum. Brian E. Hayden, Derek Pletcher, Jens-Peter Suchsland and Laura J. Williams. *Physical Chemistry Chemical Physics*, 11, 1564–1570, 2009.
DOI: 10.1039/b817553e
- Synthesis and application of RuSe_{2+δ} nanotubes as a methanol tolerant electrocatalyst for the oxygen reduction reaction. Pedro H. C. Camargo, Zhenmeng Peng, Xianmao Lu, Hong Yang and Younan Xia. *Journal of Materials Chemistry*, 19, 1024–1030, 2009.
DOI: 10.1039/b816565c
- The influence of support and particle size on the platinum catalysed oxygen reduction reaction. Brian E. Hayden, Derek Pletcher, Jens-Peter Suchsland and Laura J. Williams. *Journal of Materials Chemistry*, 19, 1024–1030, 2009.
DOI: 10.1039/b816565c
- A new application for nickel foam in alkaline fuel cells. F. Bidault, D.J.L. Brett, P.H. Middleton, N. Absond, N.P. Brandon. *International Journal Hydrogen Energy*, 34, 6799–6808, 2009. DOI: 10.1016/j.ijhydene.2009.06.035
- An improved cathode for alkaline fuel cells F. Bidault, D.J.L. Brett, P.H. Middleton, N. Abson, and N.P. Brandon. *International Journal Hydrogen Energy*, 35, 1783–1788, 2010. DOI: 10.1016/j.ijhydene.2009.12.035
- Electrocatalytic Properties of Pt Nanowires Supported on Pt and W Gauzes. Eric P. Lee, Zhenmeng Peng, Wei Chen, Shaowei Chen, Hong Yang, and Younan Xia. *ACS Nano*, 2, 2167–2173, 2008. DOI: 10.1021/nn800458p
- Direct Oxidation of Methanol on Pt Nanostructures Supported on Electrospun Nanofibers of Anatase. Eric Formo, Zhenmeng Peng, Eric Lee, Xianmao Lu, Hong Yang, and Younan Xia. *Journal of Physical Chemistry. C*, 112, 9970–9975, 2008. DOI: 10.1021/jp803763q
- CO Oxidation on Gold in Acidic Environments: Particle Size and Substrate Effects. Brian E. Hayden, Derek Pletcher, Michael E. Rendall, and Jens-Peter Suchsland. *Journal of Physical Chemistry. C*, 111, 17044–17051, 2007.
DOI: 10.1021/jp074651u
- Homogenization of the current density in polymer electrolyte fuel cells by in-plane cathode catalyst gradients. M. Santis, S.A. Freunberger, A. Reiner, F.N. Büchi. *Electrochimica Acta* 51, 5383–5393, 2006.
DOI: 10.1016/j.electacta.2006.02.008
- The evolution of the performance of alkaline fuel cells with circulating electrolyte. P. Gouérec, L. Poletto, J. Denizot, E. Sanchez-Cortezon, J.H. Miners. *Journal of Power Sources* 129, 193–204, 2004.
DOI: 10.1016/j.jpowsour.2003.11.032